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Introduction to Ionic Bonding and

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Chapter 8 Covalent

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Chemistry Chapter 8: Covalent  
Bonding. Covalent Bond. Molecule.  
Diatomic molecule. Molecular  
compound. A bond formed by the  
sharing of electrons between atoms.  
A neutral group of atoms joined  
together by covalent bonds. A  
molecule consisting of two atoms.

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8.1 Molecular Compounds > 21  
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The molecular structure of water  
shows how the oxygen ...

How are atoms joined together to  
make compounds with ...  
It will categorically ease you to look  
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## Pearson Chemistry Chapter 8

Vocabulary. Covalent bond. Molecule.

Diatomic molecule. Molecular compound. bond formed by the sharing of electrons between atoms. a neutral group of atoms joined together by covalent bonds. a molecule that contains two atoms. a compound composed of molecules.

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Flashcards and Study Sets ...

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Covalent Bonding •In covalent bonds, atoms share electrons.

•There are several electrostatic interactions in these bonds:

Attractions between electrons and nuclei, Repulsions between electrons, Repulsions between nuclei. Chapter 8

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Chapter Review Answers  
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Reserved. In general, covalent  
bonding results from an imbalance  
between the attractions and  
repulsions of the nuclei and electrons  
involved. • The nuclei and electrons  
attract each other. • Nuclei repel  
other nuclei.

Covalent Bonding - Scarsdale Public  
Schools  
Chapter 8 Covalent Bonding 183  
Section Review Objectives • State a  
rule that usually tells how many  
electrons are shared to form a  
covalent bond • Describe how

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electron dot formulas are used • Predict when two atoms are likely to be joined by a double or a triple covalent bond • Distinguish between a single covalent bond and other covalent bonds

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8.4 Polar Bonds and Molecules > -  
Pittsfield High School  
Chapter 8 Answer Key Covalent  
Bonding Pearson | [www.dougnukem.com](http://www.dougnukem.com)  
Covalent bonding is a form of  
chemical bonding between two non  
metallic atoms which is characterized  
by the sharing of pairs of electrons  
between

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Education, Inc. 106. The Octet Rule<sup>4</sup>.

- Atoms tend to gain, lose, or share electrons until they are surrounded by eight valence electrons; this is known as the octet rule.
- An octet consists of full sand psubshells.
- We know that  $ns^2np^6$  is a noble gas configuration.

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