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Revision | Class 12

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~~Properties 13.1~~

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Properties, the

van't Hoff factor,

and Molality

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Forces and Boiling  
Points Colligative  
Properties

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Properties calculate  
all of them! Worked  
out problem(s).

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\u0026 Raoult's

law | 12 Chemistry |

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Molality?!?

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~~03 | Relative~~

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~~Pressure due to~~

~~Non Volatile Solute~~

~~Colligative~~

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A we have And  
discussed, solutions  
have different  
properties than  
either the solutes  
or the solvent used  
to make the  
solution. Those  
properties can be  
divided into two  
main  
groups--colligative  
and non-colligative  
properties.

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Colligative And  
Colligative  
Properties  
Colligative properties depend only on the number of dissolved particles in solution and not on their identity. Non-colligative properties depend on the identity of the dissolved species and the solvent.



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Colligative  
Properties of  
Solutions:

Colligative ...

Colligative  
properties depend  
only on the number  
of dissolved  
particles (that  
is—the  
concentration), not  
their identity.

Raoult's law is

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Solutions with the  
concerned with the  
vapor pressure  
depression of  
Colligative  
Properties. The  
solutions. The  
boiling points of  
solutions are  
always higher, and  
the freezing points  
of solutions are  
always lower, than  
those of the pure  
solvent.

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## 11.6: Colligative Properties of Solutions - Chemistry ...

Different Types of  
Colligative  
Properties of  
Solution Lowering  
of Vapour Pressure.  
In a pure solvent,  
the entire surface is  
occupied by the  
molecules of the  
solvent. If a...

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Elevation in Boiling Point. The boiling point of a liquid is the temperature at which the vapour pressure is equal to... ..

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Properties -  
Definition, Types,  
Examples ...  
Solutions &

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**Solutions And Colligative Properties**  
properties content :  
Types of Solutions;  
Concentration of  
soln of solids in  
liquids; Solid  
Solutions;  
Colligative  
properties;  
Lowering of vapour  
pressure; Elevation  
of boiling point;  
Depression of  
freezing point;

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Osmotic pressure;  
Molecular masses  
and colligative  
properties;

Abnormal molecular  
mass; Van't Hoff  
factor

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Colligative  
properties MCQs  
for Mht-cet 2020  
Colligative

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Properties of  
Solutions  
Colligative  
Properties of  
Solutions Depends  
on concentration of  
dissolved particles:  
doesn ' t mean if  
they are small or  
large or charge  
molecules, just the  
number of particles  
per solution. There  
are four properties.

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Properties of  
Solutions -

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Properties. The  
properties of the  
solutions which  
depend only on the  
number of solute  
particles but not on  
the nature of the



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Solute are called

Colligative  
properties. The  
four important

colligative

properties are: (i)

Relative lowering in  
vapour pressure (ii)

Elevation in boiling  
point (iii)

Depression in  
freezing point (iv)

Osmotic pressure.

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Properties |  
Chemistry, Class  
12, Solutions

There are a few solution properties, however, that depend only upon the total concentration of solute species, regardless of their identities. These

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colligative And  
properties include  
vapor pressure  
lowering, boiling  
point elevation,  
freezing point  
depression, and  
osmotic pressure.

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11.4: Colligative  
Properties -  
Chemistry  
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When  $\text{CH}_3\text{OH}$  is dissolved in water, how many particles are in solution?

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Quizizz  
Solutions colligative  
properties -  
Chemistry test 1 )  
Molarity of a

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Solutions And

expressed as: a)

the number of  
moles of a solute

present in one litre  
of the solution. b)

the number of  
moles of a solute  
present in 1000 gm  
of the solvent.

---

Solutions colligative  
properties -

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Solutions And  
Chemistry test

Colligative  
properties arise  
from the fact that  
solute affects the  
concentration of  
solvent.

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Solutions,  
Solubility, and  
Colligative  
Properties ...  
Colligative

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Solutions And

Definition.

Colligative

properties are

properties of

solutions that

depend on the

number of particles

in a volume of

solvent (the

concentration) and

not on the mass or

identity of the

solute particles.



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Colligative And  
properties are also  
affected by  
temperature.

Calculation of the  
properties only  
works perfectly for  
ideal solutions.

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Definition and  
Examples of  
Colligative  
Properties

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There are a few solution properties, however, that depend only upon the total concentration of solute species, regardless of their identities. These colligative properties include vapor pressure lowering, boiling point elevation,

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freezing point  
depression, and  
osmotic pressure.

This small set of  
properties is of  
central importance  
to many ...

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11.4 Colligative  
Properties –  
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Colligative  
properties of

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Solutions are colligative properties that depend upon the concentration of solute molecules or ions, but not upon the identity of the solute. Colligative properties include vapor pressure lowering, boiling point elevation, freezing point depression, and

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osmotic pressure.

Lowering the Vapor  
Pressure:

Properties

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Colligative  
Properties -  
Chemistry &  
Biochemistry  
In chemistry,  
colligative  
properties are  
those properties of  
solutions that

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Solutions And Colligative Properties

depend on the ratio of the number of solute particles to the number of solvent molecules in a solution, and not on the nature of the chemical species present. The number ratio can be related to the various units for concentration of a solution, for

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example, molarity, molality, normality, etc. The assumption that solution properties are independent of nature of solute particles is exact only for ideal solutions, and is ap

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properties -

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properties depend  
only on the number  
of dissolved

particles (that is,  
the concentration),  
not their identity.

Raoult ' s law is  
concerned with the  
vapour pressure  
depression of  
solutions. The  
boiling points of



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Solutions are always higher, and the freezing points of solutions are always lower, than those of the pure solvent.

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Properties of  
Solutions –  
Introductory ...  
By definition, one of

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Solutions And Colligative Properties

the properties of a solution is a colligative property if it depends only on the ratio of the number of particles of solute and solvent in the solution, not the identity of the solute. Very few of the physical properties of a solution are

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Properties - Purdue  
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The properties of  
dilute solutions  
which depend only  
on number particles  
of solute present in  
the solution and not  
on their identity are

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called colligative properties (denoting depending upon collection). We shall assume here that the solute is non volatile, so it does not contribute to the vapour.

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Properties